

Thermochemistry Question Answer Guide

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Q=m.c. T. Q glass =1000.0.2. (90-20)=14000 cal. Q water =1000.1. (90-20)=70000 cal. Q calorimeter =70000 + 14000= 84000 cal. 1 mol S is 32 g. Molar combustion enthalpy of S is 84000 cal or 84 kcal. Since it is combustion enthalpy; H CombustionS = -84 kcal/mol. Thermochemistry Exams and Problem Solutions< Prev.

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Thermochemistry Thermodynamics and Energy and Temperature Thermochemistry is study of changes in energy (heat) associated with physical or chemical changes. Force = push F= m a (mass x acceleration) force units: N (newton) = kg m s-2 Work = force x distance W = F d energy units: J (joule) = kg m2 s-2

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H = M x C x T. Where M is the mass or volume of the solution being used to test the energy change, C is the Specific Heat Capacity of the solution, (for example the specific heat capacity of water is 4.1855 J g-1 K-1), and T is the change in temperature.

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Thermochemistry is a branch of Thermodynamics, a scientific field that comprises the relationships between heat, work, and other forms of energy resulting from different chemical and physical processes. Thermochemistry itself is defined as energy changes when a chemical reaction takes place. Energy is the capacity to supply heat or do work.

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The answer is given per mole of reaction. fHo (298 K) for SF 4 (g), H 2 O (l), SO 2 (g) and HF (g) are –763, –286, –297 and –273 kJ mol –1, respectively. What is the value of rHo (298 K) for the hydrolysis shown above? For this process, cHo (298 K) = –3953 kJ mol –1.

~~Chapter 2: Thermochemistry~~

42. Given that heat neutralization of strong acid and strong base is –13.7 kcal. The heat produced when one mole of HCl is mixed with 0.5 mole of NaOH will be ____