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Chemistry Chapter 12 Stoichiometry Test Answers

Remember it is a MC test, use the answers ... Practice Test Ch3 Stoichiometry (page 2 of 2) 19. The mass of element X found in 1.00 mole of each of four different compounds is 28.0 g, 42.0 g, 56.0 g, and 70 g, respectively. The possible atomic weight of X is a. 8.00 b. 14.0

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Answers: Moles and Stoichiometry Practice Problems 1) How many moles of sodium atoms correspond to 1.56x1021 atoms of sodium? 1.56 -x 1021 atoms Na x 1 mol Na = 2.59 x 10 3 mol Na 236.022 x 10 atoms Na 2) Determine the mass in grams of each of the following: a. 1.35 mol of Fe 1.35 mol Fe x 55.845 g Fe = 75.4 g Fe 1 mol Fe b. 24.5 mol O

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Chapter 9 Stoichiometry Test Answers

SOLUTION: You are askedto calculate the amount (in moles) of oxygen atoms in a given sample. You are given the mass of the oxygen sample. Use the molar mass of oxygen (1 mol O = 16.00 g) to create a conversion factor that converts mass (in grams) to amount (in mol) of oxygen. 124 g O × 1 mol O 16.00 g = 7.75 mol O.

Chapter 3 Stoichiometry

It is readily available in pdf, ppt, word, rar, txt, kindle, and also zip. Answer Key Mole/Stoichiometry.Test.Review. Answer Key Mole/Stoichiometry.Test.Review 1. 6.022x1023particles ((atoms, (molecules) ((2. 1mole (=6.022x1023particles ((1mole=molar (mass (1mole=22.4L (3.

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