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Answers: Moles and Stoichiometry
Practice Problems 1) How many moles of
sodium atoms correspond to 1.56×10^{21}
atoms of sodium? 1.56×10^{21} atoms Na
 $\times 1 \text{ mol Na} = 2.59 \times 10^{-3} \text{ mol Na}$ 236.022
 $\times 10$ atoms Na 2) Determine the mass in

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grams of each of the following: a. 1.35 mol of Fe $1.35 \text{ mol Fe} \times 55.845 \text{ g Fe} = 75.4 \text{ g Fe}$ 1 mol Fe b. 24.5 mol O

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SOLUTION: You are asked to calculate the amount (in moles) of oxygen atoms in a given sample. You are given the mass of the oxygen sample. Use the molar mass of

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oxygen (1 mol O = 16.00 g) to create a conversion factor that converts mass (in grams) to amount (in mol) of oxygen. $124 \text{ g O} \times \frac{1 \text{ mol O}}{16.00 \text{ g}} = 7.75 \text{ mol O}$.

Chapter 3 Stoichiometry

It is readily available in pdf, ppt, word, rar, txt, kindle, and also zip. Answer Key Mole/Stoichiometry.Test.Review. Answer Key Mole/Stoichiometry.Test.Review 1. 6.022×10^{23} particles ((atoms, (molecules) ((2. 1 mole (= 6.022×10^{23} particles ((1 mole = molar (mass (1 mole = 22.4L (3.

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