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S, IE2 = 2251 kJ/mol. Cl, IE2 = 2297 kJ/mol. Ar, IE2 = 2666 kJ/mol. K, IE2 = 3051 kJ/mol. Ca, IE2 = 1145 kJ/mol. Sc, IE2 = 1235 kJ/mol. Ti, IE2 = 1310 kJ/mol. It is given that the second ionization energy of Sulfur (S) is 2251 kJ/mol and that the second ionization energy of Titanium (Ti) is 1310 kJ/mol. Because the ionization energies generally increase across a period, we can find that the second ionization energy of Chlorine (Cl) and the second ionization energy of Argon (Ar) must be ...

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