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CHOICE QUESTIONS** Select the one
best answer for each question. 1. In the
reaction $\text{CaCO}_3 + 2\text{HNO}_3 \rightarrow \text{Ca}(\text{NO}_3)_2 +$

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$\text{CO}_2 + \text{H}_2\text{O}$, how many milliliters of
0.1015 M HNO_3 are required to react with
0.9857 grams of CaCO_3 ? (a) 0.1940

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QUESTIONS Select the one best answer for each question. 1. In the reaction

$\text{CaCO}_3 + 2\text{HNO}_3 \rightarrow \text{Ca}(\text{NO}_3)_2 + \text{CO}_2 + \text{H}_2\text{O}$, how many milliliters of 0.1015 M HNO_3 are required to react with 0.9857 grams of CaCO_3 ? (a) 0.1940 mL (b) 97.04 mL (c) 194.1 mL 2.

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Multiple choice questions are fundamental
survey questions which provides
respondents with multiple answer options.

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CHEM 201 Self Quiz – 2 (Gravimetric analysis/Volumetric analysis) Answer key

1. Calculate the solubility product for SrF_2 if the molar concentration of a saturated solution is $8.6 \times 10^{-4} \text{ M}$. $K_{sp} = [\text{Sr}^{2+}][\text{F}^-]^2$
 $[\text{Sr}^{2+}] = [\text{SrF}_2] = 8.6 \times 10^{-4}$
 $[\text{F}^-] = 2[\text{SrF}_2] = 2 \times 8.6 \times 10^{-4} = 17.2 \times 10^{-4} \text{ M}$

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$s_p = 8.6 \times 10^{-4}$, $17.2 \times 10^{-4} = 1.5 \times 10^{-10}$ 2.

Calcium fluoride, CaF_2 , is a sparingly ...

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