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Work is found the first law as $w = q - \Delta U$; thus $q = \Delta H$; $w = \Delta H - \Delta U$
HPVL; 4. Isothermal Process
Because U is a function only of T for an ideal gas, $\Delta U = \Delta H = 0$ for

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an isothermal process. These results also follow from the general results by using $D(T) = D(PV) = 0$ for an isothermal process.

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GASKELL Problem 1.1* The plot
of $V = V(P, T)$ for a gas is shown
in Fig. 1.1. Determine the
expressions of the two second
derivatives of the volume of this
plot. (note: the principle
curvatures of the surface are
proportional to these second
derivatives).

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Work is found the first law as $w = q - \Delta U$; thus $q = \Delta H$; $w = \Delta H - \Delta U$; 4. Isothermal Process
Because U is a function only of T for an ideal gas, $\Delta U = \Delta H = 0$ for an isothermal process. These results also follow from the general results by using $\Delta T = 0$; $\Delta(PV) = 0$ for an isothermal process.

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The isothermal expansion is
conducted at 300 K. 3.4 $H =$
42750 J, $S = 59.7 \text{ J/K}$ 3.5 The
final temperature is 323.32 K,
which is greater than 323 K
because the heat capacity
increases with increasing
temperature.

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Heat Capacities The generic heat capacity $C_p = a + b T$
 $\text{Å Å Å Å Å Å Å Å Å Å } 103 + c 105$
 $\text{Å Å Å Å Å Å Å Å Å Å Å Å Å Å Å } T^2$
 ; The heat capacities of various elements and compounds are
 $C_{pAgS} = C_p \hat{e} . 8a \text{ Ø } 21.30, b \text{ Ø } 8.54, c \text{ Ø } 1.51 <$
 $C_{pAgI} = C_p \hat{e} . 8a \text{ Ø } 30.50, b \text{ Ø } 0, c \text{ Ø } 0 <$
 $C_{pAl} = C_p + 20.75 T^2 \dots$

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Thermodynamics of Materials www.eng.utah.edu/~mse5032/gaskell.pdf · PDF file The use of P and T as the independent variables is simply a matter of choice and is done usually because P and T are easy to control and measure. MSE 3050 - University of Virginia

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