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2.12 Let p_3 , ρ_3 , and T_3 denote the conditions at the beginning of combustion, and p_4 , ρ_4 , and T_4 denote conditions at the end of combustion. Since the volume is constant, and the mass of the gas is constant, then $p_4 = \rho_4 = 11.3 \text{ kg/m}^3$. Thus, from the equation of state, $p_4 = \rho_4 RT_4 = (11.3)(287)(4000) = 1.3 \times 10^7 \text{ N/m}^2$

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