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# **Chemistry Neutralization Guided And Study Workbook Answers**

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**Lesson: Acid-Base Neutralization  
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Acid and Base Neutralization Reactions,  
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And Study

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Study Workbook Answers The full ionic

equation for the neutralization of

hydrochloric acid by sodium hydroxide is

written as follows: (4.4.2)  $\text{H}^+ + \text{Cl}^-$

$+ \text{Na}^+ + \text{OH}^- \rightarrow \text{Na}^+ + \text{Cl}^-$

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 $\text{H}^+(\text{aq}) + \text{Cl}^-(\text{aq}) + \text{H}_2\text{O}(\text{l})$  Since the acid and base are both strong, they are fully ionized and so are written as ions, as is the NaCl formed as a product.

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If 1.85 grams of  $\text{Mg}(\text{OH})_2$  reacts with 3.71...

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4.4: Neutralization Reactions - Chemistry

LibreTexts  $\text{pH} = \text{pK}_a + \log \left[ \frac{\text{mmol Base}}{\text{mmol Acid}} \right]$

$\text{pH} = \text{pK}_a + \log$

$\left[ \frac{0.65 \text{ mmol}}{0.65 \text{ mmol}} \right]$   $\text{pH} = \text{pK}_a + \log (1)$

Therefore, when the weak acid is 50% neutralized,  $\text{pH} = \text{pK}_a$ . Step 3: Solve for the pH at the equivalence point.

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equation for the neutralization of  
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written as follows:  $(4.4.2) \text{H}^+ + (\text{aq}) + \text{Cl}^-$   
 $(\text{aq}) + \text{Na}^+ + (\text{aq}) + \text{OH}^- (\text{aq}) \rightarrow \text{Na}^+ + (\text{aq})$   
 $+ \text{Cl}^- (\text{aq}) + \text{H}_2\text{O} (\text{l})$  Since the acid  
and base are both strong, they are fully  
ionized and so ...

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In chemistry, neutralization or neutralisation (see spelling differences) is a chemical reaction in which acid and a base react quantitatively with each other. In a reaction in water, neutralization results in there being no excess of hydrogen or hydroxide ions present in the solution. The pH of the neutralized solution depends on the acid strength of the reactants.

Neutralization (chemistry) - Wikipedia  
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