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1) What volume would be occupied by 100 g of oxygen gas at a pressure of 1.50 atm and a temperature of 25 °C? 2) An air-filled balloon as a volume of 225 L at 0.94 atm and 25 °C.

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Chemistry 13 4 Practice Problems Answer Key 13.4 Practice Problems
Name _____ PV = n RT R = 0.0821 atm-L/mol-K P = pressure in atm V =
volume in L n = moles T = temperature in Kelvin 1) What volume would
be occupied by 100 g of oxygen gas at a pressure of 1.50 atm and a
temperature of 25 °C? 2) An air-filled balloon as a volume

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Download 13 4 Practice Problems Chemistry Answers Problems 1. What
volume would be occupied by 100. g of oxygen gas at a pressure of
1.50 atm and a of 25°C? $0.0821 \cdot (273K)$ 2. An air-filled
balloon has a volume of 5 L at 0.94 atm and 25°C. Soon after, the
pressure changes to 0.99 atm and the temperature changes to 0°C. What is
the new volume of the

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13-4 Practice Problems 1. What volume would be occupied by 100-g of oxygen gas at a pressure of 1.50-atm and a temperature of 25 °C? 6. How many grams of argon would it take to fill a light bulb with a volume of 0.475-L at STP?

~~13-2 Practice Problems — Steffin~~

HOT! 13-4 Practice Problems Answers Chemistry - Most Popular The mole fraction in Figure 13.4.4 is that of the dissolved solute, and so there is a limit. As we shall see in the next section, if the solute is volatile,

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~~13.5: Solution Concentration—Mass Percent—Chemistry ...~~

Problem A hydrogen gas thermometer is found to have a volume of 100.0 cm³ when placed in an ice-water bath at 0°C. When the same thermometer is immersed in boiling liquid chlorine, the volume of hydrogen at the same pressure is found to be 87.2 cm³.

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