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$[4 \times 10^{-10}][3 \times 4 \times 10^{-10}]^3 \rightarrow 6.9 \times 10^{-34}$

Q: What are the concentrations of each of the ions in a saturated solution of

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PbI₂, given that the K_{sp} of PbI₂ is 8.5×10^{-9} at 25degC. If 5g pbl2 are dissolved in water to make 1L of solution at 25degC, would it be sat, unsat, or supersat

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coefficients of a balanced chemical equation. See Section 9.2 (Page 252) .

ANSWER: Correct

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Question 2 Part A

Iron reacts with oxygen in the air to produce rust (Fe_2O_3) according to the reaction shown

below. $4\text{Fe} + 3\text{O}_2 \rightarrow 2$

Fe_2O_3 How many

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moles of rust can be formed if 1.8 mols of iron are present? Hint 1.

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gram of burning
wood is

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approximately the same for both a match and a bonfire.

This is an intensive property and depends on the material (wood). However, the overall amount of produced heat depends on the amount of material; this is an extensive property.

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1. $[\text{Cr}(\text{NH}_3)_4\text{Cl}_2]$

Br is a co-ordination

compound. (March –

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9.1. Write the formulas for the following coordination compounds: (i) Tetraamminediaquacobalt(III) chloride.
(ii) Potassium tetracyanonickelate(II)
(iii) Tris

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(ethane-1,2-diamine) chromium (III) chloride. (iv) Amminebromidochloridonitrito-N- platinum (II) (v) Dichloridobis (ethane-1,2-diamine) platinum (IV) nitrate.

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Quantities. 1.

Although we define mass as the “ amount of matter in a substance, ” the units in which we measure mass are a human invention.

Atoms and molecules react on an individual particle- by-particle basis, and we have to count individual particles when doing

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Chemistry calculations.

2. Answers

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Quantities

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As we noted in

Chapter 9.2 the heat released by a reaction

carried out at

constant volume is

identical to the

change in internal

energy (ΔE) rather

than the enthalpy

change (ΔH); ΔE is

related to ΔH by an

expression that

depends on the

change in the number

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of moles of gas during the reaction.

The difference between the heat flow measured at constant volume and the enthalpy change is usually quite small, however (on the order of a few percent).

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