

Chapter 25 Nuclear Chemistry Practice Problems Answer Key

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~~Alpha Decay, Beta Decay, Gamma Decay - Electron Capture, Positron Production - Nuclear Chemistry Chapter 21 - Nuclear Chemistry: Part 3 of 9~~

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~~CHAPTER 25 NUCLEAR CHEMISTRY PRACTICE EXAMPLES 1A (E) A has a mass number of zero and an "atomic number" of 1. Emission of this electron has the effect of transforming a neutron into a proton. ${}_{94}^{241}\text{Pu} \rightarrow {}_{95}^{241}\text{Am} + {}_0^0\text{e}$ 1B (E) ${}_{58}^{58}\text{Ni} \rightarrow {}_{28}^{58}\text{Ni} + {}_{+1}^0\text{e}$ 58 Ni has a mass number of 58 and an atomic number of 28. A positron has a mass~~

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SECTION 25.1 NUCLEAR RADIATION 1. What happens to the mass number and atomic number of an atom that undergoes beta decay? 2.

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25. Terms in this set (37) Neutron Ejection. when a neutron is emitted from the nucleus. ${}^1_0\text{n}$. Particle for Neutron Ejection. ${}^4_2\text{He} + {}^1_0\text{n} + {}^4_2\text{He}$.

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Chapter 25: Nuclear Chemistry 1197 (c) We now determine the number of atoms remaining after 16 days. Because two half-lives elapse in 16 days, the number of atoms has been halved twice, to one-fourth (25%) the original number of atoms.

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