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Calculate the following
values without using a
calculator. 4.0 a. The $[H^+]$
in a solution is 1×10^{-4} M.

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Calculate the pH. 1 10 13 M

b. The pH of a solution is
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the following questions in
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Ration Ph
a. The pH of a solution is 1.0. Calculate the concentration of the solution in M.
b. The pH of a solution is 13.0. Calculate the concentration of the solution in M.

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Base Titration and pH. SHORT ANSWER . Answer the following questions in the space provided. 1. Calculate the following values without using a calculator. 4.0 . a. The [H. 3. ... MIXED REVIEW continued . 4. In a

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titration, a 25.0 mL sample
of 0.150 M .

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CHAPTER 14 REVIEW Acids and
Bases SECTION 1 SHORT ANSWER

Answer the following
questions in the space
provided. 1. Name the
following compounds as
acids: sulfuric acid a. H_2SO_4
sulfurous acid b. H_2SO_3

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Reaction Ph
2SO₃ hydrosulfuric acid c.
H₂S perchloric acid d. HClO
4 hydrocyanic acid e.
hydrogen cyanide 2.

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Bases Mixed Answers*

A Lewis acid is a species

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that accepts an electron pair to form a covalent bond. H is a Lewis acid because it can accept an electron pair from a base. For example, $\text{H}^+(\text{aq}) + \text{OH}^-(\text{aq}) \rightarrow \text{H}_2\text{O}(\text{l})$. 3. Identify the Brønsted-Lowry acid and the

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Brønsted-Lowry base on the reactant side of each of the

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yet many people who along
with don't as soon as

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reading. This is a problem.
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others to start reading, it
will be better.

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acid disappear with the

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addition of an equivalent amount of a base. It could also be said that “neutralization” of the base occurs when these two substances react to produce salts and water. 5. Bases conduct electric current.

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Like acids, bases form ions in aqueous solutions and are thus electrolytes. 458

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Reaction Ph
using a calculator. 4.0 . a.
The $[H_3O^+]$ in a solution
is 1×10^{-4} M.
Calculate the pH.

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Answer the following
questions in the space
provided. . HClO .

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hydrofluoric acid . H. 2C204

_____ acetic acid. d. Name
the acid that is present in
vinegar. 2. Answer the
following questions
according to the Brensted-
Lowry acid-base theory.

Consult Table 6 on page 485

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