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Chapter 13

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III**

CH 13 CHEMISTRY
GAS LAWS UNITS
AND CONVERSIONS

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Ideal Gas

*Problems: Crash
Course Chemistry
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LAW GAY
LUSSAC Ideal Gas
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Problems CH 13~~

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(density)

Boyle's Law

Demonstrations

Ideal Gas Law

Home Experiment

How to Use the

Ideal Gas Law in

Two Easy Steps

Chemistry 7.4d

Combined Gas Law

What are the Gas

Laws? Part 1 CH

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~~LAW'S BOYLE'S LAW~~

~~Kinetic
Molecular Theory
and the Ideal
Gas Laws Solving
Combined Gas Law
Problems~~

~~Charles' Law,
Boyle's Law,
Lussac's Law~~

~~Boyle's Law
Chapter 13 -
Properties of
Solutions: Part~~

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1 of 11 Lecture
12: Chapter 13
Properties of
Solutions -1

**Chapter 13 Gas
Laws and Kinetic
Theory The Ideal
Gas Law: Crash
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~~Chemistry Mr Z~~

AP Chemistry

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Lesson 1:

Solutions,

Solubility and

Saturation

Kinetic Theory

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Solutions |

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11 Chapter-13
Sound 11 Ncert
Solutions
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Answers

the passage of
gas particles
through a tiny
hole or orifice,

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lighter gases
diffuse more
rapidly than
heavier gases
Graham's Law the
effusion rate of
a gas is
inversely
proportional to
the square root
of its molar
mass effusion
rate of gas
a/effusion rate

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of gas b =
 $\sqrt{\frac{\text{molar mass of gas b}}{\text{molar mass of gas a}}}$
put bigger
number on top

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Chapter 13
(gases)
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Gases Vocab.

Barometer.

Millimeters of
Mercury. Torr.

Standard

Atmosphere. a
device for
measuring

atmospheric
pressure. (mm

Hg) a unit of
measurement for
pressure, also
called torr.

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another name for millimeters of mercury. a unit of measurement for pressure equal to 760 mm Hg.

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(Edition) answers
to Chapter 13 -
States of Matter

- 13.1 The
Nature of Gases
- Sample Problem
13.1 - Page 422
1 including work
step by step
written by
community
members like
you. Textbook
Authors:

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ISBN-13: 978-0-1
3252-576-3,
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- 13.1 The
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Solutions Manual

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Chemistry:
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Change • Chapter
13 255 CHAPTER
13 SOLUTIONS
MANUAL Assume
that the amount
of gas is
constant in the
following
problems. 11. A
sample of air in
a syringe exerts
a pressure of

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1.02 atm at
22.0°C. The
syringe is
placed in a
boiling water-
bath at 100.0°C.
The pressure is

GasesGases -

Weebly

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Chapter 13 Gases
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Answers points.

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Comprehending as
without
difficulty as
deal even more
than
supplementary
will allow each
success.
bordering to,
the revelation
as capably as
insight of this
chapter 13 gases
chemistry

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Answers can be
taken as
skillfully as
picked to act.
Page 2/8

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Chapter 13 –
Gases 195
Exercise 13.3 –
Equation
Stoichiometry:

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Iron is combined with carbon in a series of reactions to

form pig iron, which is about 4.3% carbon. $2C + O_2 \rightarrow 2CO$ $Fe_2O_3 + 3CO \rightarrow 2Fe + 3CO_2$ $2CO + C \rightarrow C_2O_2$ (in iron) CO_2

Pig iron is easier to shape than pure iron, and the presence

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of carbon lowers
its melting
point

Answers

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Gases - An
Introduction to
Chemistry

Chapter 13 Gases
483 t's Monday
morning, and
Lilia is walking
out of the
chemistry

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building,
thinking about
the introductory
lecture on gases
that her
instructor just
presented. Dr.
Scanlon
challenged the
class to try to
visualize gases
in terms of the
model she
described, so

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Lilia looks at her hand and tries to picture the particles in the air

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13 Gases

Chemistry

Answers Goals

and

Introductions

Section 13.1

Gases and Their
Properties Goals

To describe the
particle nature
of both real and
ideal gases. To
describe the
properties of

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gases that can be used to explain their characteristics: volume, number of particles, temperature, and pressure.

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ruyenyy.com
13.1 The Gas

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Laws. • Absolute zero is zero on the Kelvin scale. •

Charles's law states that the volume of a given amount of gas is directly proportional to its kelvin temperature at constant pressure.

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SECTION. 13.1

The Gas Laws

Charles's Law.

(cont.) Gay-

Lussac's Law.

Chemistry:

Matter and

Change

Chapter 13 Gases

1. Solids and

liquids have

essentially

fixed volumes

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and are not able to be compressed easily. Gases have volumes that depend on their conditions, and can be compressed or expanded by changes in those conditions.

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Gas Francis Howell
High School

Chapter 13:
States of Matter
- Chemistry by
Anna Chapter
13 "States of
Matter" 2.

Section 13.1 The
Nature of Gases

OBJECTIVES:

Describe the
assumptions of
the "kinetic

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theory" as it
applies to
gases. 3.

Chemistry - Chp
13 - States of
Matter It will
enormously ease
you to look
guide chemistry
chapter 13
states of matter
as you such as.

Chemistry

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Chapter 13

States Of Matter

Study Guide

Answers

Chapter 13 Study
Guide Answers 1.

Describe the assumptions/postulates of the kinetic-molecular theory of gases:

a. Gases are composed of tiny particles in

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constant rapid,
 (“random” or “st
raight-lined”?)
motion. b. Gases
are separated by
relatively huge
distances. The
volume of the
particles is
essentially
zero. c.

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Redlands Unified
School

1 CK-12

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States of Matter

13.1 Kinetic-

Molecular Theory

Practice

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Change companion course helps students learn the essential chemistry lessons of gases.

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Change Chapter
13: Gases ...
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Standard Review
Worksheet. 1.

While the
barometer is
used to measure
atmospheric
pressure, a
device called a
mercury
manometer is
used to measure
the pressure of
samples of gas
in the

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Laboratory. A manometer consists basically of a U-shaped tube filled with mercury, with one arm of the U open to the atmosphere.

Chapter 13:
Standard Review
Worksheet

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Chapter 13 – Gases and Pressure Gases in the

Atmosphere • The atmosphere of Earth is a layer of gases surrounding the planet that is retained by Earth's gravity.

- By volume, dry air is 78%

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nitrogen, 21%

oxygen, 0.9%

argon, 0.04% CO

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