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Electron Configurations Part 1- Electrons  
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C 1s 2s 2p 3s

As an example, the ground state  
configuration of the sodium atom is 1s 2  
2s 2 2p 6 3s 1, as deduced from the  
Aufbau principle (see below). The first

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excited state is obtained by promoting a 3s electron to the 3p orbital, to obtain the 1s 2s 2 2p 6 3p 1 configuration, abbreviated as the 3p level. Atoms can move from one configuration to ...

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## Electron configuration - Wikipedia

This decides the electron capacity of the shells. The K shell contains a 1s subshell hence it can carry 2 electrons, the L shell has 2s and 2p, and can carry 8 electrons. The M shell contains 3s, 3p, and 3d, and can carry 18 electrons. The N shell containing 4s, 4d, 4p and 4f, can carry 32 electrons.

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## Electron Configuration Chart for All Elements in the ...

However there are numerous exceptions; for example the lightest exception is

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chromium, which would be predicted to have the configuration  $1s^2 2s^2 2p^6 3s^2 3p^6 3d^4 4s^2$ , written as  $[\text{Ar}] 3d^4 4s^2$ , but whose actual configuration given in the table below is  $[\text{Ar}] 3d^5 4s^1$ .

---

Electron configurations of the elements  
(data page ...

An atom has the electron configuration  $1s^2 2s^2 2p^6 3s^2 3p^5$ . The electron dot symbol for this element is? a. X surrounded by seven dots. b. X surrounded by five dots. c. X surrounded by two dots. d. X surrounded by three dots. Explanation please?? :) Answer Save. 1 Answer. Relevance. Genuine. Lv 5. 8 years ago.

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An atom has the electron configuration  $1s^2 2s^2 2p^6 3s^2$  ...

A.  $1s^2 2s^1$ . B.  $1s^2 2s^2 2p^5$ . C.  $1s^2$

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$2s^2 2p^6 3s^2$ . D.  $1s^2 2s^2 2p^6 3s^2 3p^1$ . Answer Save. 1 Answer. Relevance. William. 4 years ago. Favorite Answer. B (fluorine) B needs to gain one more electron to complete its p orbitals and thus its second shell. It can do this by taking an electron from sodium (forming an ionic bond). Sodium will ...

---

If sodium (Na) has an electron configuration of  $1s^2 2s^2$  ...

A.  $1s^2 2s^2 2p^6 3s^2$  B.  $1s^2 2s^2 2p^6 3s^2 3d^4$  C.  $1s^2 2s^2$  ...” in ? Chemistry if there is no answer or all answers are wrong, use a search bar and try to find the answer among similar questions.

---

Which of the following is a reasonable ground-state ...

The electron configuration of an atom is

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1s 2s<sup>2</sup>2p<sup>6</sup>3s<sup>2</sup>3p<sup>4</sup>". The number of unpaired electrons in this atom are b. 3 c. 5 d. no correct answer given 22. The correct electron sublevel structure for 25Mn is a. 1s 2s 2p 3s 3d b. 1s 2s22p<sup>6</sup>2a<sup>3</sup>3s 3p c. 1s22s<sup>2</sup>2p 3s 3p 4s24p d. 1s 2s 2p 3s 3p 4s 3d 23.

---

Solved: A. 2 21. The Electron Configuration Of An Atom Is ...

As orbitals correspond to number of the subshell. 1 corresponds to s orbital. 2 corresponds to 2s 2p. 3 corresponds to 3s 3p 3d. 4 corresponds to 4s 4p 4d 4f. Thus 1s 2s 2p 3s 3p 3d the next will be 4s.... 297 views View 1 Upvoter

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What is the next atomic orbital in the series 1s, 2s, 2p ...

a. 1s 2s 2p 3s 3p c. 1s 2s 2p 3s 3p 4s 4p b.

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1s 2s 2p 3s 3p 4s d. 1s 2s 2p 3s 3p 3d. 1s 2s 2p 3s 3p. What is the charge on the strontium ion? a. 2- c. 1 b. 1- d. 2. 1-The octet rule states that, in chemical compounds, atoms tend to have \_\_\_\_\_. a. the electron configuration of a noble gas

---

## Chemistry test chapter 7 You'll Remember | Quizlet

An illustration of the shape and relative size of 1s, 2s and 2p orbitals. Click the check boxes to show and hide the atomic orbitals. Explore other atomic orbitals. s-orbitals | p-orbitals | d-orbitals. 4.7 (30)

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Atomic Orbitals - shape and relative size of 1s, 2s and 2p ...



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Which are impossible? (a)  $1s^2 2s^2 3s^2$  (b)  $1s^2 2p^3$  (c)  $1s^2 2s^3 2p^5$  (d)  $1s^2 2s^2 2p^7$  (e)  $1s^2 2s^2 2p^6 3s^1$  (f)  $1s^2 2s^2 2p^6 3s^2 3d^1$ . Buy Find arrow\_forward.  
Chemistry: Principles and Reactions. 8th Edition. William L. Masterton + 1 other. Publisher: Cengage Learning. ISBN: 9781305079373.

---

Which of the following electron configurations ( a ? f ...

(c)  $1s^2 2s^2 2p^3 3s^3 3p^4$  (d)  $1s^2 2s^2 2p^3 3s^3 3p^4 5^9$ .

(5 points) Please select statements that satisfy the Pauli exclusion principle. a) Electron state can hold no more than two electrons. b) Electrons with the same state must have opposite spins. c) Electron state can hold no more than four electrons.

---

Solved: (c)  $1s^2 2s^2 2p^3 3s^3 3p^4$  (d)  $1s^2 2s^2 2p^3 3s^3 3p^4 5^9$

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3p<sup>4</sup> 5 9. (5 Poin ...

Mg - 1s<sup>2</sup> 2s<sup>2</sup> 2p<sup>6</sup> 3s<sup>2</sup> b. F - 1s<sup>2</sup> 2s<sup>2</sup>

2p<sup>5</sup> c. Si - 1s<sup>2</sup> 2s<sup>2</sup> 2p<sup>6</sup> 3s<sup>2</sup> 3p<sup>2</sup> d.

Al- 1s<sup>2</sup> 2s<sup>2</sup> 2p<sup>6</sup> 3s<sup>2</sup> 3p<sup>1</sup> e. P - 1s<sup>2</sup>

2s<sup>2</sup> 2p<sup>6</sup> 3s<sup>2</sup> 3p<sup>3</sup> f. Cl - 1s<sup>2</sup> 2s<sup>2</sup>

2p<sup>6</sup> 3s<sup>2</sup> 3p<sup>5</sup> 2. Which of the following

electron configurations correspond to

ground states (lowest energy) and which

correspond to excited states? ...

---

Chemistry HW6 - sas.upenn.edu

2p 5 b) 1s 2 2s 1 c) 1s 2 2s 2 2p 6 d) 1s 2 s

2 2p 6 3s 2 3p 5 e) 1s 2 2s 2 2p 6 3s 2 3p 6

4s 1 f) 1s 2 2s 2 2p 6 3s 2 3p 6 4s 2 3d 10

4p 6 11. Specify the group of the periodic

table in which each of the following

elements is found: a) [Ne]3s 1 b) [Ne]3s 2

3p 3 c) [Ne]3s 2 3p 6 d) [Ar]4s 2 3d 8 12.

Arrange the following atoms in order of ...

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2p 5 b 1s 2 2s 1 c 1s 2 2s 2 2p 6 d 1s 2 s 2  
2p 6 3s 2 3p ...

2s shields the atom better than 2p because the s orbitals is much closer and surrounds the nucleus more than the p orbitals, which extend farther out. 3p shields better than 3d, because p orbitals are closer to the nucleus than the 3d orbitals.

---

## Penetration & Shielding - Chemistry

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$$\begin{aligned} \text{(a) N} &: 1s^2 2s^2 2p^3 \\ \text{(b) Si} &: 1s^2 2s^2 2p^6 3s^2 3p^2 \\ \text{(c) Fe} &: 1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^6 \\ \text{(d) Te} &: 1s^2 2s^2 2p^6 3s^2 3p \dots \end{aligned}$$

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Using complete subshell notation (1s 2s 2p 6 , and so ...

2 2p 6 3s 1 1s 2 2s 2 2p 6 3s 2 3p 1 1s 2 2s

# Download File PDF C 1s 2s 2p 3s 3p 4s C

2 2p 6 3s 2 3p 3 1. Examine the boarding house diagrams in Model 1. Match each symbol below with its most likely meaning. \_\_\_\_ a. I. Bunk bed for boarders \_\_\_\_ b. II. Manager's code for the number of boarders in the house and their room assignments. \_\_\_\_ c. 1s 2 2s 2 2p 6 3s 1 III. Boarder Sunny Rooms 2 POGIL™ Activities for High School ...

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2 2p 6 3s 1 1s 2 2s 2 2p 6 3s 2 3p 1 1s 2 2s 2 2p 6 3s 2 ...

b. c. d. c. 1s 2s 2p 3s 3p 3d 4s! 1s 2s 2p 3s 3p 4s 3d 1s 2s 2p 3s 3p 3d 1s 2s 2p 3s 3p 3d 1s 2s 2p 3s 3p 4s-4d 3. . What is the symbol of the neutral atom with the following electron orbital diagram?

---

Solved: 1. Which Of The Following Electron ... - Chegg.com

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There are 3 electrons in 2p orbitals. Total number of p electrons in N=3 b) The electronic configuration of Si :  $1s^2 2s^2 2p^2 3s^2 3p^2$  There are two 1s electrons, two 2s electrons, and two 3s electrons. Total number of s electrons in Si =  $2 + 2 + 2 = 6$  c) The electronic configuration of S :  $1s^2 2s^2 2p^6 3s^2 3p^4$  The 3d orbitals are empty.

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