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faculty consultants during the AP Reading in June 1999. The actual scores that these students received, as well as a brief explanation of why, are indicated. Commentary on Free-Response Question 1 This was a typical first-year college chemistry question dealing with ionic equilibria in aqueous solution. The

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Questions and Answers of 1999 AP Chemistry Multiple Choice. Terms in this set (75) (A) Activation energy (B) Free energy (C) Ionization energy (D) Kinetic energy (E) Lattice energy 1. The energy required to convert a ground-state atom in the gas phase to a gaseous positive ion. C (Ionization Energy)

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Answers to the 1999 AP Chemistry Examination 1999 AP Chemistry MC Exam - (Discussion of answers) Written by Zach Jones 1. C This is the definition of ionization energy. Notice that I.E. is not the energy to make any type of ion. Its only the energy required to lose electrons to become a cation. Your CP teacher probably didn't tell you that the

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Questions and Answers of 1999 AP Chemistry Multiple Choice. Terms in this set (75) (A) Activation energy (B) Free energy (C) Ionization energy (D) Kinetic energy (E) Lattice energy 1.

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oval on the answer sheet. Give only one answer to each question. If you change an answer, be sure that the previous mark is erased completely. Example: Chicago is a (A) state (B) city (C) country (D) continent (E) village Sample Answer Many candidates wonder whether or not to guess the answers to questions about which they are not certain.

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1999 AP Chemistry Scoring Guidelines - College Board. if answers are correct with different units and units are explicitly indicated (e.g., for part (ii), 4.02×10^{22} kJ is acceptable) (b) (i) Energy is emitted. 1 pt The $n = 6$ state is at a higher energy than the $n = 2$ state.

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1999 AP Chemistry free response answers? 1) The longest wavelength of light with enough energy to break the Cl-Cl bond in $\text{Cl}_2(\text{g})$ is 495 nm. Calculate the minimum energy in kJ/mol of the Cl-Cl bond.

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